

Thermochemistry Question Answer Guide

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$Q = m \cdot c \cdot \Delta T$. $Q_{\text{glass}} = 1000.0 \cdot 2 \cdot (90 - 20) = 14000 \text{ cal}$. $Q_{\text{water}} = 1000.1 \cdot (90 - 20) = 70000 \text{ cal}$. $Q_{\text{calorimeter}} = 70000 + 14000 = 84000 \text{ cal}$. 1 mol S is 32 g. Molar combustion enthalpy of S is 84000 cal or 84 kcal. Since it is combustion enthalpy; $\Delta H_{\text{Combustion S}} = -84 \text{ kcal/mol}$. Thermochemistry Exams and Problem Solutions < Prev.

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Thermochemistry Thermochemistry and Energy and Temperature Thermochemistry is study of changes in energy (heat) associated with physical or chemical changes. Force = push $F = m a$ (mass x acceleration) force units: N (newton) = kg m s^{-2} Work = force x distance $W = F d$ energy units: J (joule) = $\text{kg m}^2 \text{s}^{-2}$

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$\Delta H = M \times C \times -\Delta T$. Where M is the mass or volume of the solution being used to test the energy change, C is the Specific Heat Capacity of the solution, (for example the specific heat capacity of water is $4.1855 \text{ J g}^{-1} \text{ K}^{-1}$), and ΔT is the change in temperature.

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Thermochemistry is a branch of Thermodynamics, a scientific field that comprises the relationships between heat, work, and other forms of energy resulting from different chemical and physical processes. Thermochemistry itself is defined as energy changes when a chemical reaction takes place. Energy is the capacity to supply heat or do work.

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The answer is given per mole of reaction. $\Delta_f H^\circ$ (298 K) for SF_4 (g), H_2O (l), SO_2 (g) and HF (g) are -763, -286, -297 and -273 kJ mol^{-1} , respectively. What is the value of $\Delta_r H^\circ$ (298 K) for the hydrolysis shown above? For this process, $\Delta_c H^\circ$ (298 K) = -3953 kJ mol^{-1} .

~~Chapter 2: Thermochemistry~~

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42. Given that heat neutralization of strong acid and strong base is -13.7 kcal. The heat produced when one mole of HCl is mixed with 0.5 mole of NaOH will be _____

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